

We can define the electrochemical potential as:  

$$\mu_{1} = \mu_{0+} RTLn\alpha_{1} + zF\psi_{1}$$

$$\mu_{2} = \mu_{0+} RTLn\alpha_{2} + zF\psi_{2}$$
At equilibrium  $\mu_{1} = \mu_{2}$ 

$$\therefore RTLn\alpha_{1} + zF\psi_{1} = RTLn\alpha_{2} + zF\psi_{2}$$

$$\psi_{1-}\psi_{2} = E = \frac{RT}{zF}Ln\frac{\alpha_{1}}{\alpha_{2}} = \frac{RT}{zF}Ln\frac{C_{1}}{C_{2}}$$

$$E = \frac{RT}{zF}Ln\frac{C_{1}}{C_{2}}$$
Nernst equation

























## The "sodium theory" of the action potential

Action potentials exhibit an overshoot. Thus the peak of the action potential is well above zero. Hodgkin and Katz suggested (in 1949) that this was due to a rapid and selective increase increase in the permeability towards sodium. Thus  $g_{Na}$  transiently becomes much greater than  $g_k$ . How can this idea be tested?

















